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Chemistry Live! - Worked Solutions 2 Chapter 9 - The Mole Concept 9.1 Since a sulfur atom ($A_r = 32$)

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weighs twice as much as an oxygen atom ($A_r = 16$), 32 g of sulfur atoms and 16 g of oxygen atoms both contain the same number of atoms. Answer: 32 g 9.2 (a) 1 mole of H atoms = 1 g 1 mole of H₂ molecules = $2 \times 1 = 2$ g. (b) 1 mole of N atoms = 14 g

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solutions for problems in chapter 18 is solved. 1lc; 1stp; 2lc; 2stp; 3lc; 3stp; 4lc; 4stp; 5lc; 5stp; 6lc; 6stp; 7lc; 7stp; 8lc; 8stp; 9lc; 9stp; 10lc; 11lc; 12lc; 13lc; 14lc; 15lc; 16lc; 17lc; 18lc; 19lc; 20lc; 21lc; 22lc; 23lc; 24lc; 25lc; 26lc; 27lc; 28lc; 29lc; 30lc; 31lc; 32lc; 33lc; 34lc; 35lc; 36lc; 37lc; 38lc; 39lc; 40lc; 41lc; 42lc; 43lc; 44lc; 45lc; 46lc; 47lc; 48lc; 49lc; 50lc; 51lc; 52lc; 53lc; 54a; 55a; 56a; 57a; 58a; 59a; 60a; 61a; 62a; 63a; 64a; 65a; 66a; 67a; 68a; 69a; 70a ...

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18

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Chapter 9 – The Mole Concept W9.1 (a) 1 mole of Li atoms = 7 g (b) 1 mole of Na atoms = 23 g (c) 1 mole of Ca atoms = 40 g (d) 1 mole of Fe atoms = 56 g (e) 1 mole of Ag atoms = 108 g (f) 1 mole of Pb atoms = 207 g W9.2 (a) 1 mole of Cl atoms = 35.5 g 10 moles of Cl atoms = 35.5×10

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